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Chapter 9: Chemical Names and Formulas. STUDY. PLAY. Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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To name a binary molecular compound, use the following guidelines: 1. Write the names of the elements in the order listed in the formula 2. Use prefixes appropriately to indicate the number of each kind of atom 3. End the name of the second

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element with the suffix -ide

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First, confirm that the compound is a binary molecular compound. Name the elements in the order listed in the formula. Use prefixes to indicate the number of each kind of atom. Omit the prefix -mono when the formula contains only one atom of the first element in the name. The suffix of the name of the second element is -ide.

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1. when the name of the anion (X) ends in -ide the acid name begins with the prefix hydro-. The stem of the anion has the suffix -ic and is followed by the word acid (ex. HCl= hydrochloric acid) 2. when the anion name ends in -ite the acid name is the stem of the anion with the suffix -ous

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followed by the word acid (ex. H_2SO_3 --
X=sulfite = sulfurous acid)

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Writing the Names and Formulas for
Bases. A base is a compound that
produces OH^- in water. When naming a
base, you name it like any other compound
that starts with a regular or transition
metal. Ex: NaOH = sodium hydroxide.
When writing the formula for a base,
remember to balance charges. Ex:
magnesium hydroxide = $\text{Mg}(\text{OH})_2$

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Chapter 9 Chemical Names and Formulas
83 SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

~~Name Date Class CHEMICAL NAMES AND FORMULAS 9~~
SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC COMPOUNDS
I. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c.

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b. f. 3. Name the following binary ionic compounds. sodium sulfide g. CuCl: h. SnCl_4 . a. MnO_2 c. d. SrBr_2 221 Chapter 9
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Chemical Names and Formulas 281

CHAPTER 9 Assessment 42. a. 2- b. 1+ c.

1- d. 3+ 43. a. 2+ b. 2+ c. 3+ d. 1+

44. a. barium ion b. iodide ion c. silver ion d.

mercury(II) ion 45. cyanide, CN^- and

hydroxide, OH^- 46. a. hydroxide ion b.

lead(IV) ion c. sulfate ion d. oxide ion 47.

zero; A compound is electrically neutral.

48. The symbols for the cation and

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Index Numbers. Thoroughly describes the relationships between coloring reactions."

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Principles of Organic Chemistry begins by exploring the step-by-step processes (or mechanisms) by which reactions occur to create molecular structures. It then describes some of the many ways these reactions make new compounds, examined by functional groups and corresponding common reaction mechanisms.

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